



UMC Utrecht

Matrix effecten: de achilleshiel van kwantitatieve LC-MS?

Ing. Klaas Stienstra* en Dr. Erik van Maarseveen*

Apotheek Laboratorium, Universitair Medisch Centrum
Utrecht

**beide sprekers hebben geen belangenverstremgeling te melden.*





LCMS voor de (ziekenhuis)apothekers (?)



Fundamentals of Mass Spectrometry

<https://www.youtube.com/watch?v=EzvQzImBuq8&app=desktop>

How does it work?

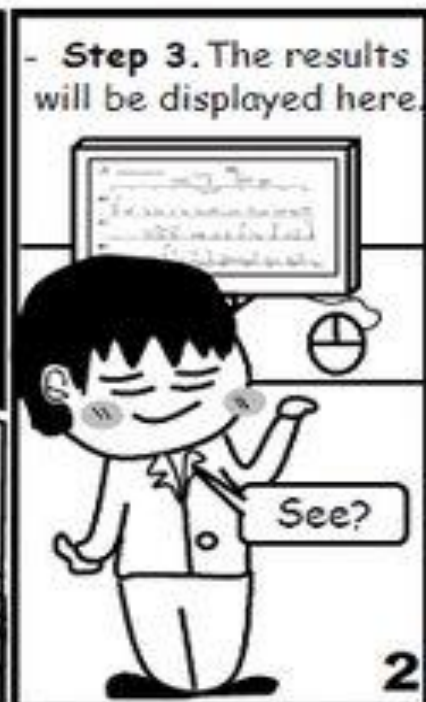
<https://www.youtube.com/watch?v=YA5DDt6bMlw>

Wonderlab

– Mass Spectrometry

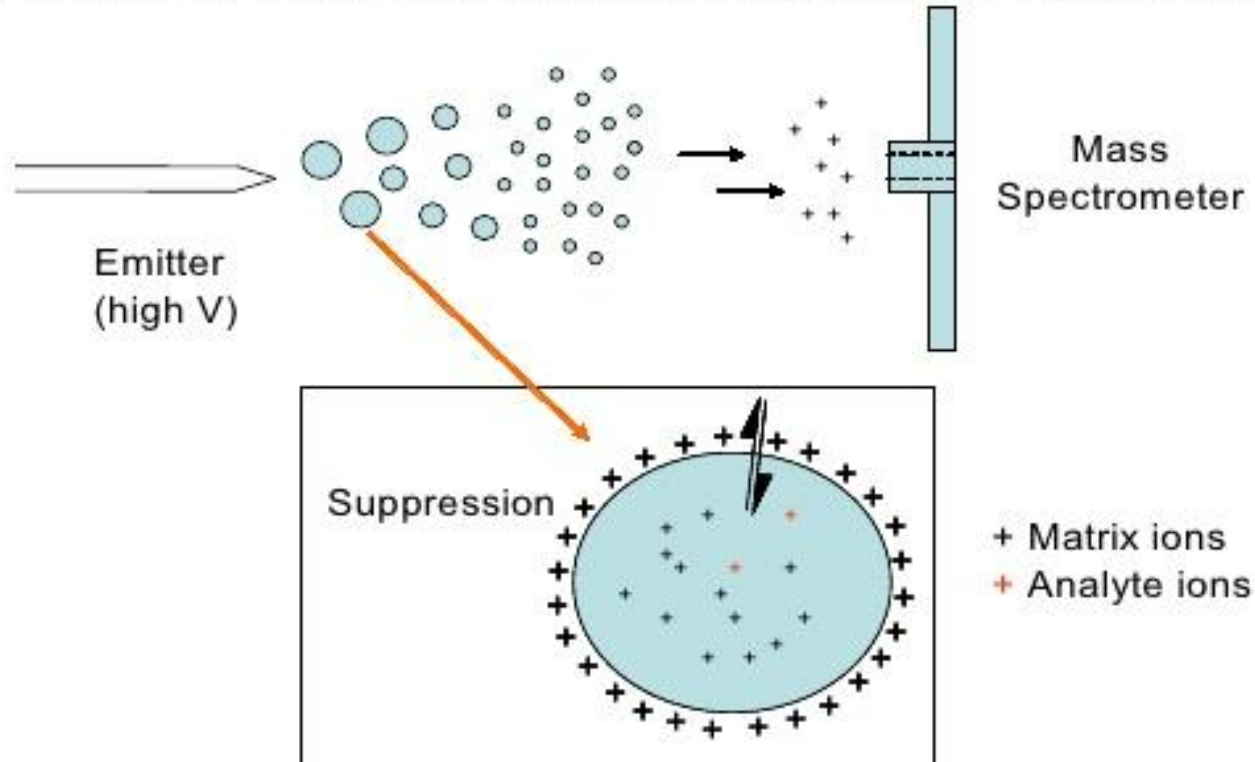


2 Weeks Ago...



Matrix Effect in LC-ESI-MS

Matrix effects are the result of competition between nonvolatile matrix components and analyte ions for access to the droplet surface for transfer to the gas phase.

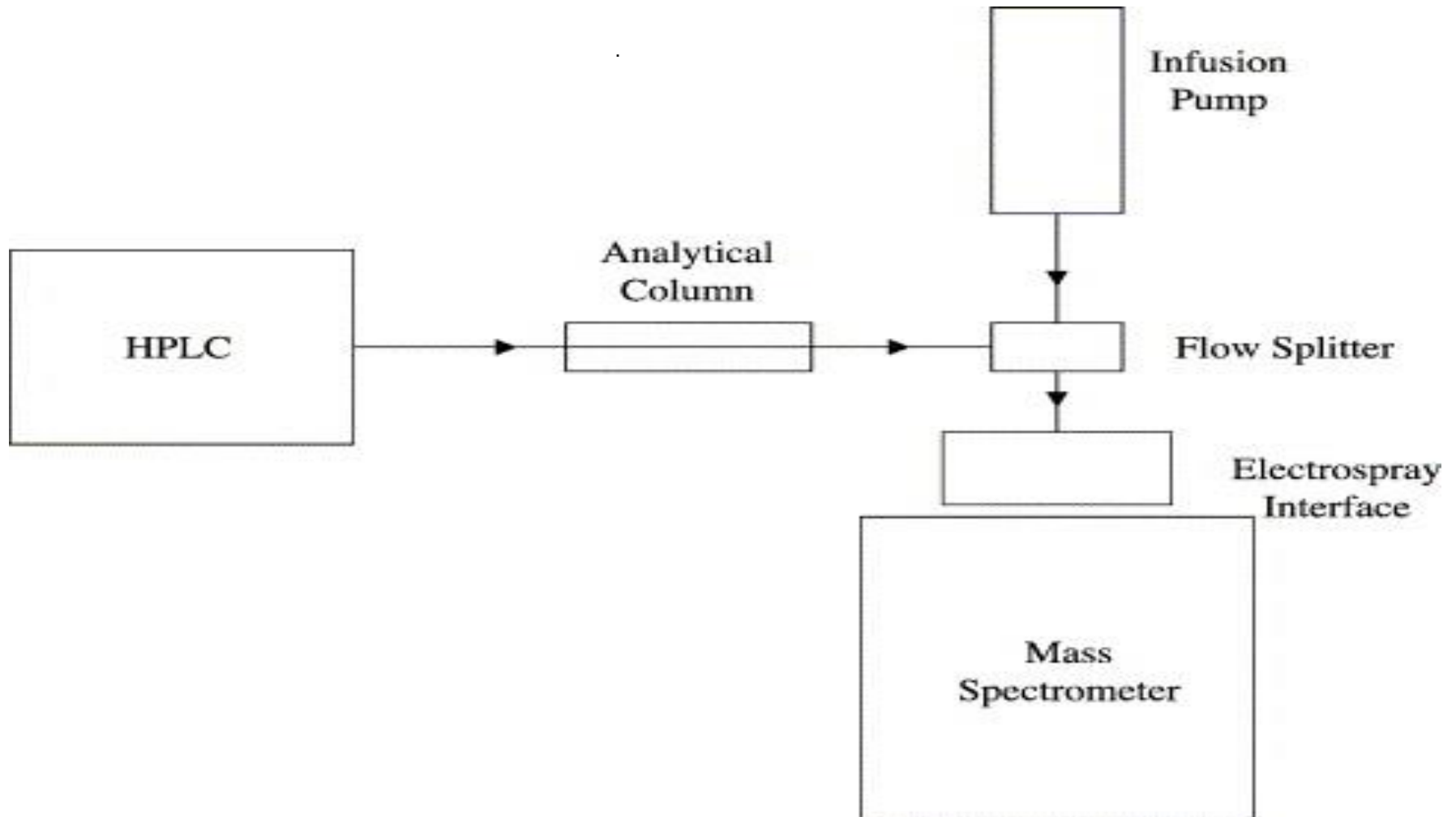


Depending on the environment in which the ionization and ion evaporation processes take place, this competition may effectively decrease (**ion suppression**) or increase (**ion enhancement**) the efficiency of formation of the desired analyte ions.

De Matrix-detective....



Schematic of the postcolumn



Paul J. Taylor

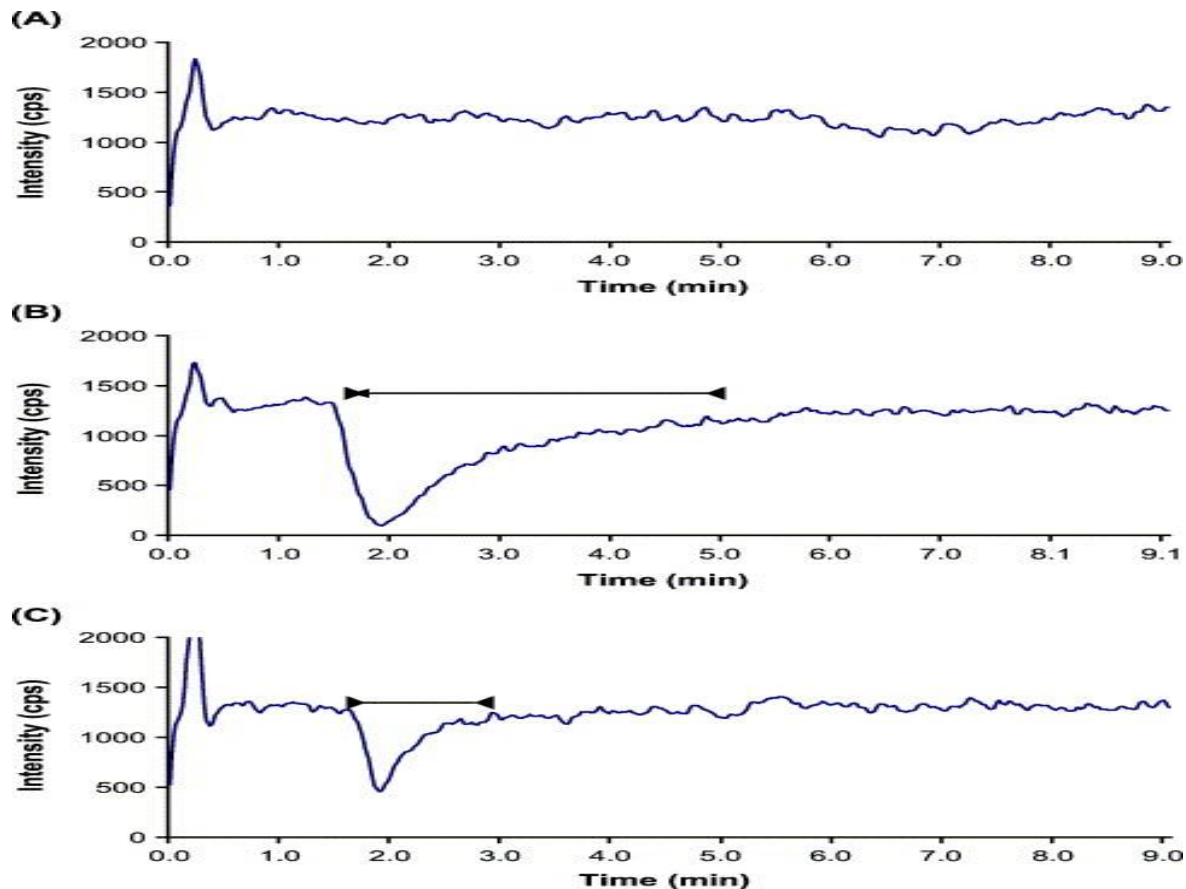
Matrix effects: the Achilles heel of quantitative high-performance liquid chromatography–electrospray–tandem mass spectrometry

Clinical Biochemistry, Volume 38, Issue 4, 2005, 328–334



UMC Utrecht

het Waaron van de Postkolom



Comparison of (A) mobile phase, (B) whole blood sample prepared by protein precipitation, and (C) a whole sample prepared by solid phase extraction [33] by the postcolumn infusion method.

Process efficiency is key!

Table 2.

Validation data of an HPLC–ESI–MS/MS method, for the measurement of cyclosporin, using the pre- and postextraction technique

Parameter ± standard deviation	Cyclosporin concentration (µg/L)			Internal standard (300 µg/L)
	30	400	1500	
Matrix effect (%)	-7.5 ± 3.6	-7.8 ± 6.1	-3.6 ± 4.1	-7.3 ± 7.5
Absolute recovery (%)	75.7 ± 3.3	82.2 ± 4.6	83.7 ± 5.0	79.3 ± 6.7
Process efficiency (%)	69.8 ± 5.1	75.6 ± 2.8	80.7 ± 5.6	73.6 ± 9.3
Intersubject variability ^a (%)	2.4	2.3	1.0	^b

Cyclosporin parameters are calculated on five measurements and internal standard parameters on 15 measurements. Methodology reported in Taylor et al. [40].

a Expressed as coefficient of variation ($n = 5$).

b Not applicable.

Matrix problemen: een voorbeeld

Casus: combi-run anti-epileptica in plasma m.b.v. LC-MS/MS



Methode 1

- 17 anti-epileptica in plasma
- Opwerking: eiwit precipitatie
- LC-MS Ultimate 3000 UHPLC met TSQ Acces MAX
 Triple quadrupole MS (Thermo Scientific)
- Ion source ESI in pos/neg mode
- Kolom Acquity UPLC BEH C18 (50x2.1 mm; 1.7 µm)
- Column temp 50°C
- Eluens A: Water + 0.1% ammoniumacetaat
 B: Methanol
- Analysetijd 9 min
- 3 isotoop gelabelde interne standaarden

Validatie volgens FDA guidelines

	<i>Felbamaat</i>	<i>Fenobarbital</i>	<i>Fenytoïne</i>
Lineariteit	14,3 – 130 mg/L	7,38 – 60,1 mg/L	3,49 – 23,3 mg/L
Nauwkeurigheid LLQ	WR: 2,2 %	WR: 4,5 %	WR: 2,4%
	BR: 9,0 %	BR: 0,0 %	BR: 3,3 %
Nauwkeurigheid MED	WR: 2,8%	WR: 2,1 %	WR: 1,9 %
	BR: 4,9 %	BR: 0,0 %	BR: 2,5 %
Vries/ dooi effect	WR: 2,8%	WR: 2,1 %	WR: 1,9 %
	BR: 4,9 %	BR: 0,0 %	BR: 2,5 %
Autosamplerstabiliteit, t =	20 uur	20 uur	20 uur
Matrix effect	8,0 %	4,5 %	0,7 %
Specificiteit	Acc		

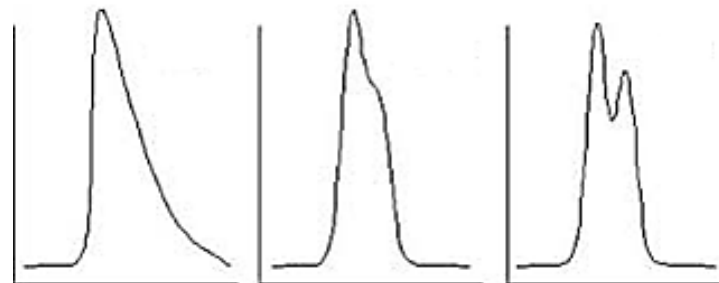
Alle componenten voldoen aan de vooraf gestelde validatie eisen

Methode gevalideerd

Problemen

Na verloop van tijd ontstonden er problemen

- Verhoging van druk
- Slechte piekvorm
- Afwijking van de QC's



Oorzaken en Oplossingen

Mogelijke oorzaken van de problemen

- **Verstopping van de kolom**
- **Kolom degradatie**
- **Schimmelgroei in leidingen UHPLC systeem en de kolom**

Oplossingen van de problemen

- **Vervangen kolom**
- **Gebruik maken van een voorkolom**
- **Monsters opwerken met proteïn precipitation plate**
- **Spoelen systeem volgens procedure leverancier**

Optimalisatie methode 2

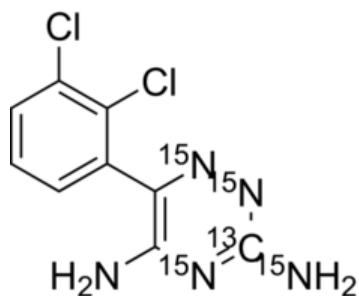
- Overstap naar andere kolom leverancier (Kinetex XB-C18 75x2,1 mm; 2,6 μ m)



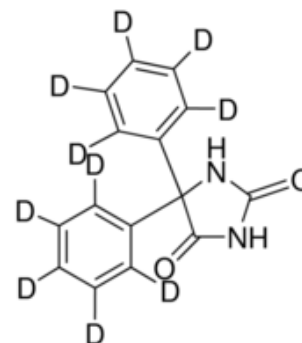
- Gradient aanpassen

Optimalisatie methode 2

- Isotoop gelabelde interne standaard voor alle componenten



Lamotrigine-13C,15N4



Fenytoine-d10

Resultaten hervalidatie

	<i>Felbamaat POS</i>	<i>Fenobarbital NEG</i>	<i>Fenytoïne NEG</i>
Lineariteit	14,0 – 129 mg/L	6,91 – 58,4 mg/L	3,47 – 24,4 mg/L
Nauwkeurigheid LLQ	WR: 10,3 % BR: 3,8 %	WR: 7,5 % BR: 8,3 %	WR: 9,0% BR: 12,1 %
Nauwkeurigheid MED	WR: 4,9% BR: 11,3 %	WR: 3,9 % BR: 0,6 %	WR: 4,0 % BR: 4,2 %
Matrix effect %	0,4 %	-3,5 %	-2,8 %
Recovery %	97 %	100 %	100 %
Specificiteit	Akk	Akk	Akk

Alle componenten voldoen aan de vooraf gestelde validatie eisen

Methode gevalideerd

Methode 1 vs Methode 2

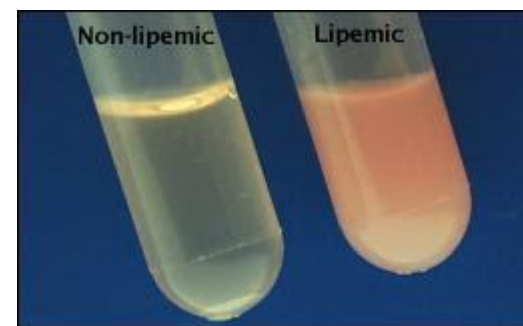
Analyseren patient plasma met methode 1 en methode 2 maar met isotoop gelabelde ISTD

Component
10-OH-Carbamazepine
Gabapentine
Lamotrigine
Levetiracetam
Topiramaat
Lacosamide
Ethosuximide
Fenobarbital
Fenytoine
Valproinezuur



Samenvatting

- Tijdens validatie verschillende matrices getest zoals Kalfs serum, KKG T serum, Biorad plasma en Patienten plasma. Hierbij geen matrix effect aangetoond!
- Leerpunt 1: Stabiele isotoop gelabelde interne standaard bood goede correctie voor hele proces
- Leerpunt 2: Gebruik van representatief patientenmateriaal



Hoe het matrix monster te temmen?



1920s, Doreen taking an alligator ride

Questions?

1. How do you know if a reaction is spontaneous?

2. How do you know if a reaction is at equilibrium?

3. How do you know if a reaction is exothermic?

4. How do you know if a reaction is endothermic?

5. How do you know if a reaction is reversible?

6. How do you know if a reaction is irreversible?

7. How do you know if a reaction is spontaneous at a given temperature?

8. How do you know if a reaction is at equilibrium at a given temperature?

9. How do you know if a reaction is exothermic at a given temperature?

10. How do you know if a reaction is endothermic at a given temperature?

11. How do you know if a reaction is reversible at a given temperature?

12. How do you know if a reaction is irreversible at a given temperature?

13. How do you know if a reaction is spontaneous at a given temperature and pressure?

14. How do you know if a reaction is at equilibrium at a given temperature and pressure?

15. How do you know if a reaction is exothermic at a given temperature and pressure?

16. How do you know if a reaction is endothermic at a given temperature and pressure?